

**Paper Reference(s) 4CH1/1C 4SD0/1C**  
**Pearson Edexcel International GCSE (9–1)**

**Chemistry**  
**UNIT: 4CH1**  
**Science (Double Award) 4CH1/4SD0**  
**PAPER: 1C**

|                    |
|--------------------|
| <b>Total Marks</b> |
|--------------------|

**Friday 17 May 2024 – Morning**

**Time: 2 hours**

**In the boxes below, write your name, centre number and candidate number.**

|                         |  |  |  |  |  |
|-------------------------|--|--|--|--|--|
| <b>Surname</b>          |  |  |  |  |  |
| <b>Other names</b>      |  |  |  |  |  |
| <b>Centre Number</b>    |  |  |  |  |  |
| <b>Candidate Number</b> |  |  |  |  |  |

## **YOU MUST HAVE**

**Calculator, ruler**

## **YOU WILL BE GIVEN**

**Diagram Booklet**

**Periodic Table**

## **INSTRUCTIONS**

**Answer ALL questions.**

**Answer the questions in the spaces provided in this Question Paper or in the separate Diagram Booklet – there may be more space than you need.**

**Show all the steps in any calculations and state the units.**

## **INFORMATION**

**The total mark for this paper is 110.**

**The marks for EACH question are shown in brackets – use this as a guide as to how much time to spend on each question.**

**There may be spare copies of some diagrams.**

**ADVICE**

**Read each question carefully before you start to answer it.**

**Write your answers neatly and in good English.**

**Try to answer every question.**

**Check your answers if you have time at the end.**

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**Answer ALL questions.**

**Some questions must be answered with a cross in a box ☐. If you change your mind about an answer, put a line through the box ☒ and then mark your new answer with a cross ☐.**

**1 The list below gives the names of some substances.**

**bromine**

**chlorine**

**diamond**

**ethene**

**iodine**

**lithium**

**methane**

**water**

**(a) Look at the table for Question 1(a) in the Diagram Booklet. Complete the table by choosing a substance from the box that matches each description.**

**Each substance may be used once, more than once or not at all.**

**(5 marks)**

**1 continued.**

**(b) Describe a test for chlorine.  
(2 marks)**

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**(Total for Question 1 = 7 marks)**

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**2 This question is about the reactivities of metals.**

- (a) Look at the table for Question 2(a) in the Diagram Booklet. The table shows the reactions of four metals, **P**, **Q**, **R** and **S**, with water and with dilute hydrochloric acid.**

**The letters are not the symbols of the elements.**

- (i) Deduce the order of reactivity of the metals.  
(1 mark)**

|                |       |
|----------------|-------|
| most reactive  | _____ |
| ↓              | _____ |
|                | _____ |
|                | _____ |
| least reactive | _____ |

- (ii) Give the letter of the metal that could be zinc.  
(1 mark)**

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**(continued on the next page)**

**2(a) continued.**

- (iii) Give a word equation for the reaction between aluminium and hydrochloric acid.  
(1 mark)**

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- (iv) Give the name of a metal that could be P.  
(1 mark)**

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- (v) Give a reason why the reaction of Q with dilute hydrochloric acid is not done.  
(1 mark)**

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**(continued on the next page)**

**2 continued.**

- (b) Look at the diagram for Question 2(b) in the Diagram Booklet. It shows the apparatus used to demonstrate the reaction between aluminium and iron(III) oxide.**

**When the magnesium fuse is lit, a very exothermic reaction occurs.**

**This is the equation for the reaction.**



- (i) State what is meant by the term EXOTHERMIC.**  
**(1 mark)**

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**(continued on the next page)**



2(b) continued.



(ii) State why aluminium displaces iron.  
(1 mark)

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(iii) Explain why this reaction is a redox reaction.  
(2 marks)

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(Total for Question 2 = 9 marks)

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- 3 Look at the diagram for Question 3 in the Diagram Booklet. It represents an atom of element **Z**.

**Z** is not the symbol of the element.

- (a) (i) Give the number of the group to which element **Z** belongs.  
(1 mark)

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- (ii) Give the number of the period to which element **Z** belongs.  
(1 mark)

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- (iii) Give the formula of the compound that forms when **Z** reacts with fluorine.  
(1 mark)

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(continued on the next page)

Turn over

**3 continued.**

**(b) One mole of  $Z$  contains  $6.0 \times 10^{23}$  atoms.**

**Calculate the number of electrons in one mole of atoms of element  $Z$ .**

**Give your answer in standard form.  
(2 marks)**

**number of electrons = \_\_\_\_\_**

**(continued on the next page)**

**3 continued.**

- (c) Look at the table for Question 3(c) in the Diagram Booklet. A sample of element **Z** contains three isotopes. The table shows the numbers of particles in the nucleus of each isotope and the percentage abundance of each isotope.**

**Use the information in the table to calculate the relative atomic mass ( $A_r$ ) of element **Z**.**

**Give your answer to one decimal place.  
(4 marks)**

$A_r =$  \_\_\_\_\_

**3 continued.**

**(d) Deduce the name of element Z.**  
**(1 mark)**

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**(Total for Question 3 = 10 marks)**

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- 4 Caffeine is a stimulant found in coffee, tea and some soft drinks.

(a) The molecular formula of caffeine is  $\text{C}_8\text{H}_{10}\text{N}_4\text{O}_2$

- (i) Determine the number of atoms in one molecule of caffeine.

(1 mark)

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- (ii) Calculate the relative formula mass ( $M_r$ ) of caffeine.

(2 marks)

$M_r =$  \_\_\_\_\_

- (iii) Give the empirical formula for caffeine.

(1 mark)

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Turn over

**4 continued.**

**(b) Look at the diagram for Question 4(b) in the Diagram Booklet. Ethanol can be obtained from a solution of caffeine in ethanol using this apparatus.**

**(i) Give the name of the method of separation shown in the diagram.  
(1 mark)**

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**(ii) Describe what happens to the ethanol vapour in apparatus X.  
(2 marks)**

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**4 continued.**

**(c) Calcium bromide is an ionic compound.**

**Look at the table for Question 4(c) in the Diagram Booklet. It shows the formulae and melting points of caffeine and calcium bromide.**

**The relative formula mass of calcium bromide is similar to the relative formula mass of caffeine.**

**Explain why calcium bromide has a much higher melting point than caffeine.**

**(5 marks)**

**Answer space continues on the next page.**

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**4(c) continued.**

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**(Total for Question 4 = 12 marks)**

**Turn over**

- 5 A student uses paper chromatography in an experiment to separate the dyes in four different felt tip pens, E, F, G and H.**

**Look at the diagram for Question 5 in the Diagram Booklet. It shows the appearance of the paper before and after the experiment.**

- (a) (i) The chromatography paper is placed in a solvent. Explain why the spots on the baseline are placed above the level of the solvent.  
(2 marks)**

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**5(a) continued.**

- (ii) Explain which two felt tip pens contain the same dye.  
(2 marks)**

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**5(a) continued.**

**(iii) The student thought that both **F** and **G** contained only one dye.**

**Explain why the student can only be certain about one of these dyes.**

**(2 marks)**

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**(continued on the next page)**

**5 continued.**

- (b) Calculate the  $R_f$  value for the dye in G.  
(3 marks)**

**Show your working.**

**$R_f$  value = \_\_\_\_\_**

**(Total for Question 5 = 9 marks)**

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**6 This question is about some Group 1 elements and their compounds.**

**(a) A teacher adds a small piece of sodium to a trough of water.**

**(i) Give two observations that are made when sodium reacts with water.  
(2 marks)**

**1** \_\_\_\_\_

\_\_\_\_\_

**2** \_\_\_\_\_

\_\_\_\_\_

**(continued on the next page)**

**6(a) continued.**

- (ii) After the reaction has stopped, the teacher adds a few drops of phenolphthalein to the solution in the trough.**

**Explain the colour of the phenolphthalein after it is added to the solution.  
(2 marks)**

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**6 continued.**

- (b) A student does a flame test to see if a white solid contains lithium ions.**

**They clean a platinum wire before using it for the flame test.**

- (i) Explain why the student needs to clean the platinum wire.  
(2 marks)**

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**(continued on the next page)**



**6(b) continued.**

**(ii) What is the colour of the flame if the solid contains lithium ions?  
(1 mark)**

- ☐ **A    lilac**
- ☐ **B    orange**
- ☐ **C    red**
- ☐ **D    yellow**

**(continued on the next page)**

6 continued.

- (c) Potassium aluminium sulfate can be used in baking.

Anhydrous potassium aluminium sulfate has the formula  $\text{KAl}(\text{SO}_4)_2$

- (i) Give the formula of each ion in potassium aluminium sulfate.  
(2 marks)

potassium ion

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aluminium ion

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sulfate ion

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(continued on the next page)

Turn over

6(c) continued.

- (ii) Potassium aluminium sulfate is normally found as a hydrated salt, with the formula  $\text{KAl}(\text{SO}_4)_2 \cdot x\text{H}_2\text{O}$

When **23.7 g** of the hydrated salt is heated to remove all the water, **12.9 g** of the anhydrous salt is formed.

Calculate the value of **X**.  
(4 marks)

[for  $\text{KAl}(\text{SO}_4)_2$ ,  $M_r = 258$       for  $\text{H}_2\text{O}$ ,  $M_r = 18$ ]

Answer space continues on the next page.

6(c)(ii) continued.

[for  $\text{KAl}(\text{SO}_4)_2$ ,  $M_r = 258$       for  $\text{H}_2\text{O}$ ,  $M_r = 18$ ]

**X** = \_\_\_\_\_

(Total for Question 6 = 13 marks)

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**7 This question is about nitrogen and its compounds.**

**(a) What is the approximate percentage by volume of nitrogen in the atmosphere?  
(1 mark)**

☐ **A 1%**

☐ **B 20%**

☐ **C 70%**

☐ **D 80%**

**(b) Look at the diagram for Question 7(b) in the Diagram Booklet. Complete the dot-and-cross diagram for a molecule of nitrogen.**

**Show outer electrons only.  
(2 marks)**

**(continued on the next page)**

**7 continued.**

**(c) Nitrogen dioxide produced in car engines reacts with water vapour and oxygen in the atmosphere to form nitric acid.**

**(i) Give a chemical equation for this reaction.  
(2 marks)**

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**(ii) Nitric acid forms acid rain.**

**State one environmental effect of acid rain.  
(1 mark)**

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7 continued.

(d) Ammonium carbonate contains nitrogen.

(i) What is the formula of ammonium carbonate?  
(1 mark)

☐ A  $\text{NH}_3\text{CO}_3$

☐ B  $(\text{NH}_3)_2\text{CO}_3$

☐ C  $\text{NH}_4\text{CO}_3$

☐ D  $(\text{NH}_4)_2\text{CO}_3$

(ii) A technician finds an unlabelled bottle on a shelf that could be ammonium carbonate solution.

Describe tests that the technician should do to confirm that the solution contains ammonium ions and carbonate ions.  
(6 marks)

Answer space continues on the next 2 pages.

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Turn over

**7(d)(ii) continued.**

[illegible]

**Turn over**



**7(d)(ii) continued.**

[illegible]

**(Total for Question 7 = 13 marks)**

8 This question is about hydrocarbons.

(a) The molecular formula  $\text{C}_4\text{H}_8$  represents all the isomers of an alkene.

(i) Explain what is meant by the term ISOMERS.  
(2 marks)

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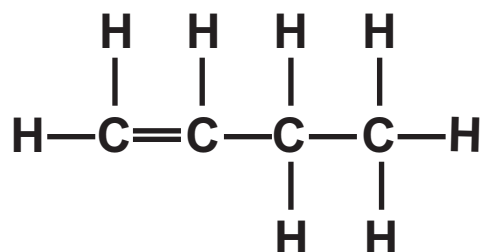
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(continued on the next page)

8(a) continued.

- (ii) The displayed formula of one of the isomers of the alkene is shown.



Look at the diagram for Question 8(a)(ii) in the Diagram Booklet. Draw displayed formulae for two other alkene isomers with the molecular formula  $\text{C}_4\text{H}_8$   
(2 marks)

- (b) **But-1-ene** reacts with bromine to form a compound with molecular formula  $\text{C}_4\text{H}_8\text{Br}_2$

What is the name given to this type of reaction?  
(1 mark)

- ☐ A addition
- ☐ B combustion
- ☐ C decomposition
- ☐ D substitution

(continued on the next page)

Turn over

8 continued.

(c) The alkene  $\text{C}_3\text{H}_6$  can be polymerised to form the polymer poly(propene).

(i) Draw the repeat unit of poly(propene).  
(1 mark)

(continued on the next page)

8(c) continued.

(ii) These are two methods for disposing of polymers such as poly(propene).

method 1 burying them in landfill sites

method 2 burning them to release heat energy

State one environmental problem linked to each of these methods of disposal.  
(2 marks)

method 1

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method 2

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(continued on the next page)

Turn over

8 continued.

- (d) Complete combustion of one mole of an alkane produces **396 g** of carbon dioxide and **180 g** of water.

This is the equation for the reaction.



Calculate the values of **X**, **y** and **Z**.  
(3 marks)

[for  $\text{CO}_2$ ,  $M_r = 44$       for  $\text{H}_2\text{O}$ ,  $M_r = 18$ ]

**X** = \_\_\_\_\_

**y** = \_\_\_\_\_

**Z** = \_\_\_\_\_

**8 continued.**

**(e) In a petrol engine, incomplete combustion occurs because there is a limited supply of oxygen.**

**(i) Petrol contains octane,  $\text{C}_8\text{H}_{18}$**

**Look at the equation for Question 8(e)(i) in the Diagram Booklet. Complete the equation for this reaction, including state symbols.**

**(2 marks)**

**(ii) Explain one problem for humans caused by a product of this incomplete combustion.**

**(2 marks)**

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**(Total for Question 8 = 15 marks)**

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- 9 Look at the diagram for Question 9(a) in the Diagram Booklet. A student uses this apparatus to investigate the rate of reaction between marble chips and dilute hydrochloric acid.

This is the equation for the reaction.



- (a) During the reaction the reading on the balance decreases because mass is lost from the flask.

- (i) State why mass is lost from the flask.  
(1 mark)

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- (ii) State the purpose of the cotton wool.  
(1 mark)

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(continued on the next page)

Turn over



**9 continued.**

**(b) Look at the graph for Question 9(b) in the Diagram Booklet. This is a graph of the student's results.**

**(i) Explain the shape of the graph.**

**You should assume that the marble chips are in excess.**

**(4 marks)**

**Answer space continues on the next page.**

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**9(b)(i) continued.**

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**(continued on the next page)**

**9(b) continued.**

- (ii) Look again at the graph for Question 9(b) in the Diagram Booklet. On the grid, draw the curve you would expect to obtain if the student uses the same volume of hydrochloric acid but with half the concentration.**

**Assume that all other conditions are kept the same.**

**(2 marks)**

**(continued on the next page)**

**9 continued.**

- (c) The student repeats the experiment using the same mass of smaller marble chips.**

**Explain, using particle collision theory, how using smaller marble chips would affect the rate of this reaction.**

**Assume that all other conditions are kept the same as in the initial experiment.**

**(3 marks)**

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**9(c) continued.**

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**(Total for Question 9 = 11 marks)**

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- 10 A student investigates the reaction between magnesium and nitric acid.**

**The student uses this method.**

- **add  $40\text{ cm}^3$  of dilute nitric acid to a glass beaker**
- **record the temperature of the acid**
- **find the mass of a strip of magnesium ribbon**
- **add the magnesium ribbon to the nitric acid**
- **when all the magnesium has reacted, record the highest temperature reached**

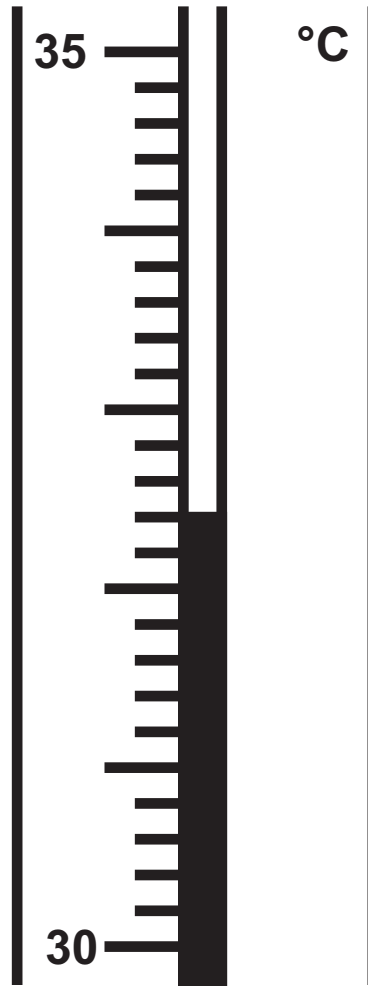
- (a) Look at the equation for Question 10(a) in the Diagram Booklet. Complete the chemical equation for this reaction.**

**(1 mark)**

**(continued on the next page)**

10 continued.

- (b) The thermometer shows the highest temperature reached.



Look at the table for Question 10(b) in the Diagram Booklet. Complete the table by giving the temperatures to the nearest 0.1 °C (2 marks)

(continued on the next page)

**10 continued.**

- (c) (i) Show that the heat energy change ( $Q$ ) for this reaction is about **2800 J**.  
(2 marks)**

**[for  $1.0 \text{ cm}^3$  of solution, mass =  $1.0 \text{ g}$ ]**

**[for the solution,  $c = 4.2 \text{ J/g/}^\circ\text{C}$ ]**

**(continued on the next page)**



10(c) continued.

- (ii) The mass of magnesium used by the student was **0.12 g**.

Calculate the value of the enthalpy change ( $\Delta H$ ), in **kJ/mol**, for the magnesium reacting with nitric acid.

Give your answer to two significant figures, including a sign in your answer.  
(4 marks)

$\Delta H =$  \_\_\_\_\_ **kJ/mol**

**10 continued.**

- (d) Explain why using a polystyrene cup, instead of a glass beaker, would give a more accurate result.  
(2 marks)**

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**(Total for Question 10 = 11 marks)**

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**TOTAL FOR PAPER = 110 MARKS**  
**END OF PAPER**